

THE REACTION OF AZIRIDINE WITH SODIUM BOROHYDRIDE

Brian P. Robinson

Department of Chemistry, Acadia University, Wolfville, N. S., Canada

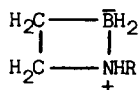
and

K. A. H. Adams

Department of Chemistry, Mount Allison University, Sackville, N. B., Canada

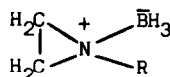
(Received in USA 7 October 1968; received in UK for publication 13 November 1968)

The reaction between aziridine and sodium borohydride was reported by Akerfeldt and Hellstrom (1) to give the four-membered heterocyclic compound (Ia) rather than aziridine-borane (IIa). This reaction was of interest to us in connection with some other work. We



Ia R = H

b R = CH₃



IIa R = H

b R = CH₃

considered that the evidence presented by Akerfeldt and Hellstrom in favour of the proposed structure Ia was inconclusive since it was based on one of several possible interpretations of an NMR spectrum. We, therefore, repeated the reaction of aziridine with sodium borohydride in aqueous tetrahydrofuran according to the reported procedure (1).

A product was isolated which had the same physical properties as those reported by Akerfeldt and Hellstrom and which after purification by vacuum sublimation analysed as C₂H₈BN (Found: C, 42.4; H, 14.0; B, 18.9; N, 24.6. Calc. for C₂H₈BN: C, 42.2; H, 14.2; B, 19.0; N, 24.6%). The infrared spectrum of this compound showed the presence of N-H⁺ (3300 and 2740 cm⁻¹) as would be expected for either Ia or IIa.

On the other hand, in the case of Ib and IIb, where R = CH₃, only Ib contains an N-H⁺ group. We, therefore, carried out the reaction under the same conditions with N-methylaziridine as the donor molecule and isolated a white solid (42%) which after vacuum sub-

limation melted at 57-58° and analysed as C₃H₁₀BN. (Found: C, 51.1; H, 13.8; B, 15.3; N, 19.8. Calc. for C₃H₁₀BN: C, 50.8; H, 14.2; B, 15.2; N, 19.8%). The infrared spectrum of this product showed no absorptions characteristic of an N-H group, which eliminates Ib as a possible structure and suggests that the compound is IIb. By analogy the compound obtained by Akerfeldt and Hellstrom would thus be IIa.

The ¹¹B NMR spectra of the compounds C₂H₈BN and C₃H₁₀BN, from aziridine and N-methylaziridine respectively, provided convincing evidence for the structures IIa and IIb. The presence of a BH₃ group in both compounds was evident from the presence of a symmetrical 1:3:3:1 quartet in the spectrum of each. The coupling constants J_{BH} of 91 and 96 cps are in agreement with the values reported for other amine-boranes (2) thus providing additional evidence for structures IIa and IIb. The ¹H NMR spectrum of the compound C₂H₈BN showed an A₂B₂ pattern due to the methylene hydrogens. This is consistent with the structure IIa, since the methylene hydrogens cis and trans to the BH₃ group are not magnetically equivalent and spin-spin coupling gives rise to the observed pattern. The NMR spectra of aziridine and various N-substituted aziridines have been reported to show an A₂B₂ pattern for the dimethylene moiety when the rate constant for the nitrogen inversion process is small compared with the chemical shift, ν_{AB} (3,4). In compound IIa an A₂B₂ pattern would be expected since there is no opportunity for nitrogen inversion. The ¹¹B and ¹H NMR data are summarized in the Table.

Burg and Good (5) obtained a product (m.p. about 40°) from the reaction of aziridine with diborane, which was not characterized but undoubtedly was also aziridine-borane (IIa).

Preliminary experiments show that methylamine, dimethylamine and trimethylamine when reacted with sodium borohydride under the same conditions as were used for aziridine and N-methylaziridine, give the corresponding amine-boranes which have been characterized by their melting points and infrared spectra, and in the case of trimethylamine-borane by comparison with an authentic sample.

Work on this reaction is continuing and details will be reported elsewhere.

TABLE
NMR Data

	^{11}B (in CHCl_3)	^1H (in CDCl_3) ^a			
	J_{BH} cps	$-\text{BH}_3$	$-(\text{CH}_2)_2^-$	$-\text{CH}_3$	J_{BH} cps
$\text{C}_2\text{H}_8\text{BN}$	91.3 ± 0.9	quartet $8.68 \tau^b$	A_2B_2 $7.88 \tau^c$	-	92
$\text{C}_3\text{H}_{10}\text{BN}$	96.3 ± 0.7	quartet $8.49 \tau^d$	$7.96 \tau^e$	singlet 7.4τ	95

^a TMS as the internal standard

^b centre of the quartet; broad bands at 11.0τ , 9.45τ , 7.9τ (obscured) and 6.4τ

^c centre of the symmetrical A_2B_2 pattern; multiplets at 7.63τ and 8.16τ

^d centre of the quartet; broad bands at 10.9τ , 9.3τ , 7.7τ (obscured) and 6.1τ

^e broadened singlet

ACKNOWLEDGEMENTS

We thank the National Research Council of Canada for financial assistance, and Dr. R. G. Kidd and Mr. D. R. Truax (University of Western Ontario) for the determination of the ^{11}B NMR spectra.

REFERENCES

1. S. Akerfeldt and M. Hellstrom, *Acta Chem. Scand.* **20**, 1418 (1966).
2. W. D. Phillips, H. C. Miller and E. L. Muetterties, *J. Am. Chem. Soc.* **81**, 4496 (1959).
3. A. T. Bottini and J. D. Roberts, *J. Am. Chem. Soc.* **78**, 5126 (1956); **80**, 5203 (1958).
4. F. A. L. Anet and J. M. Osyany, *J. Am. Chem. Soc.* **89**, 352 (1967).
5. A. B. Burg and C. D. Good, *J. Inorg. Nucl. Chem.* **2**, 237 (1956).